## DETERMINING **EMPIRICAL FORMULAS**

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日本所有自己的な結果、我们であって、大学院で勝手

Name.

What is the empirical formula (lowest whole number ratio) of the compounds below?

1. 75% carbon, 25% hydrogen	
2. 52.7% potassium, 47.3% chlorine	
3 - 22.1% duminum $25.4%$ photophotus $52.5%$ ox/dop	
3. 22.1% aluminum, 25.4% phosphorus, 52.5% oxygen	
4. 13% magnesium, 87% bromine	
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5. 32.4% sodium, 22.5% sulfur, 45.1% oxygen	
6 05 2% coppor 10 0% within 05 7% everyon 26 1% weter	
6. 25.3% copper, 12.9% sulfur, 25.7% oxygen, 36.1% water	
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## DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

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Solve the problems below.

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٦.	The empirical formula of a compound is $NO_2$ . Its molecular mass is 92 g/mol. What is its molecular formula?
2.	The empirical formula of a compound is CH <sub>2</sub> . Its molecular mass is 70 g/mol. What is its molecular formula?
3.	A compound is found to be 40.0% carbon, 6.7% hydrogen and 53.5% oxygen. Its molecular mass is 60. g/mol. What is its molecular formula?
4.	A compound is 64.9% carbon, 13.5% hydrogen and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?
5.	A compound is 54.5% carbon, 9.1% hydrogen and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?
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